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Synthesis and characterization of a new ion-imprinted polymer for the selective separation of thorium(IV) ions at high acidity

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A new ion-imprinted polymer (IIP) was synthesized by thermal copolymerization of bis(2-methacryloxyethyl) phosphate as functional ligand and ethylene glycol dimethacrylate as cross-linker in the presence of Th⁴⁺ as a template ion. The molar ratio of the functional ligand to Th⁴⁺ was optimized to be 4 by ³¹P NMR titration, elemental analysis, and EXAFS, which was verified by comparing the adsorption capacities and selectivities of IIPs with different composition. Thorium(IV) ions were leached out using 0.1 mol L⁻¹ ethylenediaminetetraacetic acid disodium salt and the prepared IIP was characterized by infra-red spectroscopy, thermogravimetric analysis and Brunauer–Emmett–Teller surface area measurement. It was found that the IIP could extract Th⁴⁺ from high acidity solution. The maximum adsorption capacity was as high as 33.3 mg g⁻¹ in 1.0 mol L⁻¹ HCl. Even in 6.0 mol L⁻¹ HCl, its adsorption capacity was still considerable. Moreover, the IIP exhibited good selectivity towards Th⁴⁺, especially in the presence of competing ions such as La³⁺, Eu³⁺, Yb³⁺, UO₂²⁺, and Fe³⁺.

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1. Introduction

Thorium is widely used in various areas, such as optics, radio, aeronautics, and the chemical industry.¹ More importantly, thorium is a promising fuel for the next generation of nuclear power plants.² In general, thorium, uranium, rare-earth elements and iron often coexist in minerals as well as in waste water.^{3,4} Because thorium has both chemical toxicity and radioactivity,⁵ it is necessary to detect and remove Th⁴⁺ in waste water. Up to now, many methods have been applied, such as liquid–liquid extraction,⁶ solid phase extraction (SPE),⁷ ion exchange,⁸ extraction chromatography,⁹ membrane separation¹⁰ and so on. Among all the methods, SPE is widely used because of flexible working conditions and simple procedures. In the field of SPE, it is most important to synthesize an adsorbent with high selectivity towards the target ion.¹¹

Ion-imprinted polymers (IIPs) are a new kind of adsorbent for SPE, and are usually prepared by the copolymerization of a functional ligand and a cross-linker in the presence of template ion.^{12–14} Because of its memory effect on the template ion, the IIP can extract the target ion with good selectivity. Therefore, IIP is often used to separate and pre-concentrate

trace metal ions.^{15,16} In the preparation of IIP, understanding the formation of complex compounds between metal ions and functional ligands is very helpful. Thus, many researchers studied the interaction between metal ion and functional ligand to guide their syntheses of IIPs.^{17–20}

With respect to the SPE of Th⁴⁺, IIPs exhibit good performances. Buyuktiryaki *et al.*²¹ prepared a kind of Th⁴⁺–IIP beads with *N*-methacryloyl-(L)-glutamic acid as functional ligand to selectively separate Th⁴⁺ from UO₂²⁺, La³⁺, and Ce³⁺. Lin and coworkers²² synthesized Th⁴⁺–IIPs on the surface of silica gel by using 1-phenyl-3-methylthio-4-cyano-5-acrylicacidcarbamoyl-pyrazole as functional ligand, which was used to pre-concentrate Th⁴⁺ before the analysis with UV-vis spectro-photometry. In the syntheses of Th⁴⁺–IIPs, many functional ligands were used, which is summarized in Table 1. It can be found that most of the monomers own carboxyl or amino groups, leading to the optimum pH value of the IIPs between 3 and 5. When the pH value was lower than 3, the adsorption capacities of these IIPs decreased dramatically. For example, the magnetic Th⁴⁺-imprinted chitosan resin prepared by Huang and his coworkers had a large adsorption capacity of 147.1 mg g⁻¹ in a solution of pH = 4, but the adsorption capacity was no more than 10% of the maximum value at the pH value of 1.²³ However, in some practical cases, it is necessary to extract Th⁴⁺ in high acidity environment.⁷ Recently, this problem began to attract attention in the field of ion-imprinting technology. Fasihi *et al.*²⁴ used vinyl sulfonate, a strong ion-exchanger functional ligand, to prepare Th⁴⁺–IIP, which could start to extract Th⁴⁺ with an extraction efficiency of 41% at pH = 1.6. By far, none of these Th⁴⁺–IIPs could be used at pH < 1. Therefore, there still remains a great challenge.

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Table 1 Summary of the adsorption properties of Th⁴⁺-IIPs prepared using different functional ligands in the literature

Functional ligand	pH range	Maximum adsorption capacity/mg g ⁻¹	Reference
<i>N</i> -Methacryloyl-(<i>L</i>)-glutamic acid	2–4	40.44 (pH = 3)	21
Chitosan-phthalate	2–4	61.3 (pH = 3)	25
<i>N</i> -(<i>o</i> -Carboxyphenyl)maleamic acid	2–5	35.9 (pH = 3)	26
1-Phenyl-3-methylthio-4-cyano-5-acrylic acid carbamoyl-pyrazole	1.5–5	64.8 (pH = 3.5)	22
Methacrylic acid	2–5	33.2 (pH = 3)	27
3-Methyl-1-phenyl-4-(<i>cis</i> -acylbutenoic acid)-2-pyrazolin-5-one	2–6	56.8 (pH = 4.5)	28
<i>N,N'</i> -Bis(3-allyl salicylidene) <i>o</i> -phenylenediamine	3–6	42.54 (pH = 4.5)	17
Thiosemicarbazide, 4-vinyl pyridine	2–5	67.28 (pH = 4)	29
Vinyl sulfonate	1.5–3.5	9.28 (pH = 3)	24
Chitosan	1–6	147.1 (pH = 4)	23

Acidic organophosphorus compounds, particularly organophosphorus monacids (such as bis(2-ethylhexyl)-phosphoric acid (HDEHP)¹⁰ and 2-ethylhexyl phosphoric acid mono-2-ethylhexyl ester³⁰), are widely used in the extraction of Th⁴⁺ and trivalent rare earths. Especially, they can extract Th⁴⁺ from 1 mol L⁻¹ H₂SO₄ solution.³¹ Our research group has ever used HDEHP to extract Th⁴⁺ by reversed micelles with high efficiency.^{32,33} Dioleoylphosphoric acid^{34,35} and 1,12-dodecanediol-*O,O'*-diphenyl phosphonic acid³⁵ were used as functional ligand to synthesize Zn²⁺-IIP, La³⁺-IIP, Ce³⁺-IIP, and Dy³⁺-IIP. Meanwhile, the Dy³⁺-IIP could be used at pH = 1.³⁴ However, to the best of our knowledge, acidic organo-phosphorus compounds have not been used in the preparation of Th⁴⁺-IIP. Moreover, with respect to organophosphorus monacids, their coordination behaviors with Th⁴⁺ are not well understood. In the literature, there appeared several complex compounds with different molar ratios of ligand to Th⁴⁺, e.g. Th(NO₃)(DEHP)₃,³² Th(NO₃)₂(DEHP)₂,³³ Th[H(DEHP)₂]₄,³⁶ and Th(NO₃)₂(DcyHPA)₂ (DcyHPA: dicyclohexylphosphinic acid).³⁷ Therefore, it is essential to determine the molar ratio of functional ligand to Th⁴⁺ when organophosphorus monacid is applied as functional ligand to synthesize Th⁴⁺-IIP.

Since organophosphorus monacids have strong affinities towards Th⁴⁺, Th⁴⁺-IIP was synthesized in this work by using bis(2-methacryloxyethyl) phosphate (BMAOP) as the functional ligand. Also, the molar ratio of functional ligand to Th⁴⁺ was optimized in detail.

2. Experimental

2.1 Reagents and instruments

Bis(2-methacryloxyethyl) phosphate (BMAOP, N/A, J&K Scientific, China), ethylene glycol dimethacrylate (EGDMA, 98%, Acros, Belgium) and dimethyl phosphate (DMP, 98%, Acros, Belgium) were used as received. Azobisisobutyronitrile (AIBN, A.R., Beijing Yili Fine Chemical Products Inc., China) was purified by recrystallization. Ultrapure water was used throughout the experiments. All other chemical reagents used in this study were of analytical grade and used without further purification.

The concentrations of all the metal ions were determined by inductively coupled plasma-atomic emission spectrometer (ICP-

AES, Leeman, USA) with relative standard deviation (RSD) below 5%. Thermo-gravimetric analysis (TGA) was carried out using a TGA-DSC-DTA (Q-600 SDT, Thermal Analysis Co., USA). The Brunauer-Emmett-Teller (BET) surface area was measured using an accelerated surface area & porosimetry system (ASAP 2010, Micrometer, USA). FT-IR spectra were recorded in the frequency range of 400–4000 cm⁻¹ using FT-IR spectrometer (Tensor 27, Bruker, Germany). ³¹P NMR spectra were obtained with phosphoric acid as external standard by using Bruker-400 MHz NMR (ARX400, Bruker, Switzerland). The elemental analysis was taken by using an Elemental Analyzer (Vario EL, Elementar Analysensysteme GmbH, Germany). A pH meter (Delta 320, Mettler-Toledo, Switzerland) was used to measure pH values.

Sample for Extended X-ray Absorption Fine Structure (EXAFS) measurement was prepared by mixing Th(NO₃)₄·4H₂O, DMP and triethylamine (Et₃N) at a molar ratio of 1 : 4 : 4 in dimethyl sulphoxide (DMSO) in a 2 mL plastic tube, and then measured directly. XAFS measurement at Th L3-edge in transmission mode was performed at the BL14W1 (ref. 38) in Shanghai Synchrotron Radiation Facility (SSRF). The electron beam energy was 3.5 GeV and the stored current was 260 mA (top-up). A 38-pole wiggler with the maximum magnetic field of 1.2 T inserted in the straight section of the storage ring was used. XAFS data were collected using a fixed-exit double-crystal Si (111) monochromator. The energy was calibrated using Zr foil. The photon flux at the sample position was 6.9 × 10¹¹ photons per second. The raw data analysis was performed using IFEFFIT software package according to the standard data analysis procedures.³⁹ The spectrum was calibrated, averaged, pre-edge background subtracted, and post-edge normalized using Athena program in IFEFFIT software package. The Fourier transformation of the *k*³-weighted EXAFS oscillations, *k*³ × χ(*k*), from *k* space to *R* space was performed over a range of 2.3–10.3 Å⁻¹ to obtain a radial distribution function. And data fitting was done by Artemis program in IFEFFIT.

2.2 Synthesis of Th⁴⁺-imprinted polymer

In a typical preparation process, 0.322 g (1.0 mmol) BMAOP and 0.101 g Et₃N (1.0 mmol) were dissolved in 2.5 mL DMSO, and stirred for 0.5 h. Then, 0.138 g Th(NO₃)₄·4H₂O (0.25 mmol) was added and stirred for 2 h. After adding 0.792 g EGDMA (4 mmol) and 10 mg AIBN, the mixture was purged with N₂ for 20 min and



polymerized in an oil bath at 60 °C for 18 h. The bulk polymer was grounded and stirred in 100 mL 0.1 mol L⁻¹ ethylenediaminetetraacetic acid disodium salt (Na₂EDTA) for several times to remove Th⁴⁺. Then, the polymer was washed with water and acetone in turn. At last, the dried polymer was sieved to get the particles between 80 and 200 mesh. The non-imprinted polymer (NIP) was prepared under similar conditions except the absence of Th(NO₃)₄·4H₂O and Et₃N.

2.3 Adsorption experiments

The adsorption capacities of IIP (NIP) were determined by stirring 0.0050 g IIP (NIP) with a 0.0100 L solution of Th(NO₃)₄ (15–60 mg L⁻¹) and HCl (10^{-3.8}–6 mol L⁻¹) at 25 °C. Then, the mixtures were centrifuged at 4000 rpm for 5 min, and the concentration of Th⁴⁺ in the solutions was measured by ICP-AES. The adsorption capacity (*Q*, mg g⁻¹) and distribution ratio (*k_d*, mL g⁻¹) was calculated according to eqn (1) and (2).

$$Q = (C_0 - C_e)V/W \quad (1)$$

$$k_d = [(C_0 - C_e)V/C_e W] \times 1000 \quad (2)$$

where *C*₀ and *C*_{*e*} (mg L⁻¹) are the concentrations of the metal ions before and after adsorption, respectively, while *V* (L) is the volume of the testing solution and *W* (g) the weight of adsorbent.

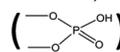
In the selectivity studies, 0.0100 g of the IIP (NIP) was shaken with a solution containing both Th⁴⁺ and the other metal ion, where the concentration of each metal ion was 20 mg L⁻¹. The selectivity coefficient *S* and the relative selectivity coefficient *k'* were calculated according to eqn (3) and (4).

$$S_{\text{Th}/M} = k_{d(\text{Th})}/k_{d(M)} \quad (3)$$

$$k' = S_{\text{IIP}}/S_{\text{NIP}} \quad (4)$$

3. Results and discussion

3.1 Optimizing the molar ratio of functional ligand to Th⁴⁺

Before the syntheses of the Th⁴⁺-IIP, the molar ratio of the functional ligand to Th⁴⁺ should be determined. The functional ligand BMAOP is not stable at room temperature and easy to polymerize. Thus, DMP was chosen as a stable substitute for BMAOP, because each of them has one phosphoric acid group () to coordinate with metal ions.

Firstly, ³¹P NMR titration was used. With the change of the concentration of DMP (deprotonated with the same amount of Et₃N), the ³¹P NMR spectra were measured at the Th(NO₃)₄ concentration of 0.25 mmol L⁻¹. As can be seen from Fig. 1, the peak of the free ligand around 0.8 ppm starts to appear when the molar ratio of DMP to Th⁴⁺ arrives at 3.9 and the intensity of the peak grows fast beyond 4, suggesting the formation of a relatively stable 4 : 1 complex compound between DMP and Th⁴⁺.

Moreover, the complex compound of DMP with Th⁴⁺ was synthesized to confirm the molar ratio of DMP to Th⁴⁺. After mixing Th(NO₃)₄, DMP and Et₃N at a molar ratio of 1 : 4 : 4 in

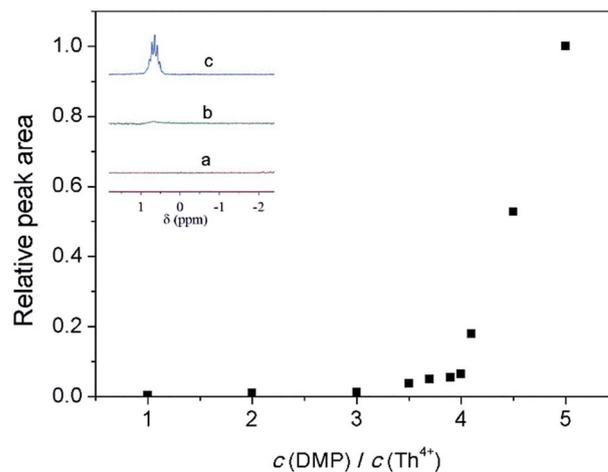


Fig. 1 ³¹P NMR study of the Th⁴⁺-DMP complex compound. The relationship between the relative peak area of free ligand and *n*(DMP)/*n*(Th⁴⁺). (The peak area of free ligand is defined as 1.0 at the *n*(DMP)/*n*(Th⁴⁺) value of 5). Inset: partial ³¹P NMR spectra of the solution with different *n*(DMP)/*n*(Th⁴⁺): 1 (a), 3.9 (b), 4.5 (c).

DMSO, water was added. Then, white precipitates were collected and dried in vacuum. Elemental analysis showed that the composition of the compound is close to that of Th(DMP)₄·H₂O (elemental analysis (%) calculated for Th(C₂H₆O₄P)₄·H₂O: C 12.81, H 3.49; found: C 12.91, H 3.34).

Furthermore, EXAFS was used to determine the coordination number of Th⁴⁺ when the molar ratio of DMP to Th⁴⁺ is 4. The fitting curve of the EXAFS data according to the single scattering theoretical phase and amplitude functions is shown in Fig. 2. The result indicates that a shell of 9.1 ± 0.7 oxygen atoms directly coordinated to Th⁴⁺ at approximately 2.39 ± 0.01 Å (Table 2). Since organophosphorus monacids often coordinate with Th⁴⁺ as bidentate ligands,³⁷ combining the results of ³¹P NMR and elemental analysis, it is reasonable to conclude that four bidentate DMPs and one monodentate ligand (water or DMSO) coordinate with Th⁴⁺ in the solution (Scheme 1).

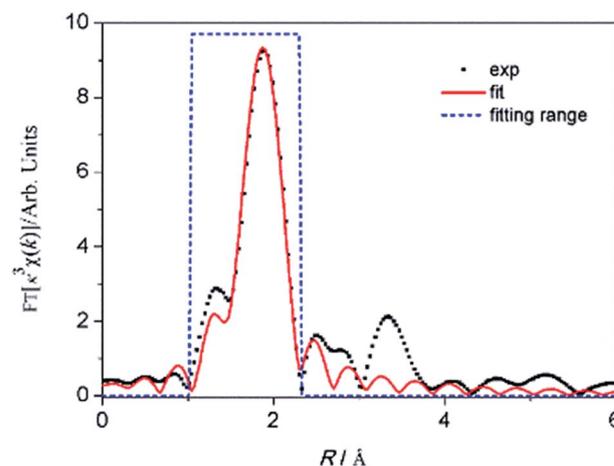
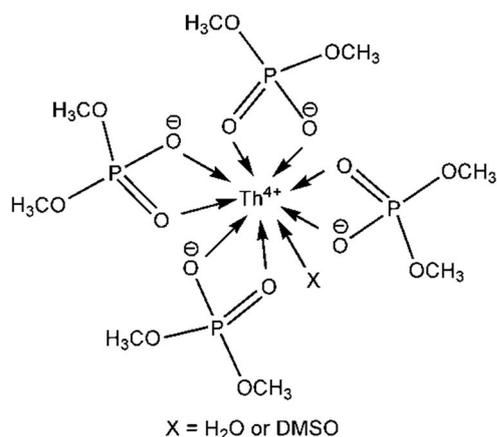


Fig. 2 The *k*³-weighted EXAFS Fourier transform spectrum (dots) and the best fitting (line) of Th⁴⁺ in the solution when *c*(DMP)/*c*(Th⁴⁺) is 4.



Table 2 Results of respective coordination number (CN) and Th^{4+} -O bond length (R) by fitting two shells of the k^3 weighted Th L3 edge EXAFS

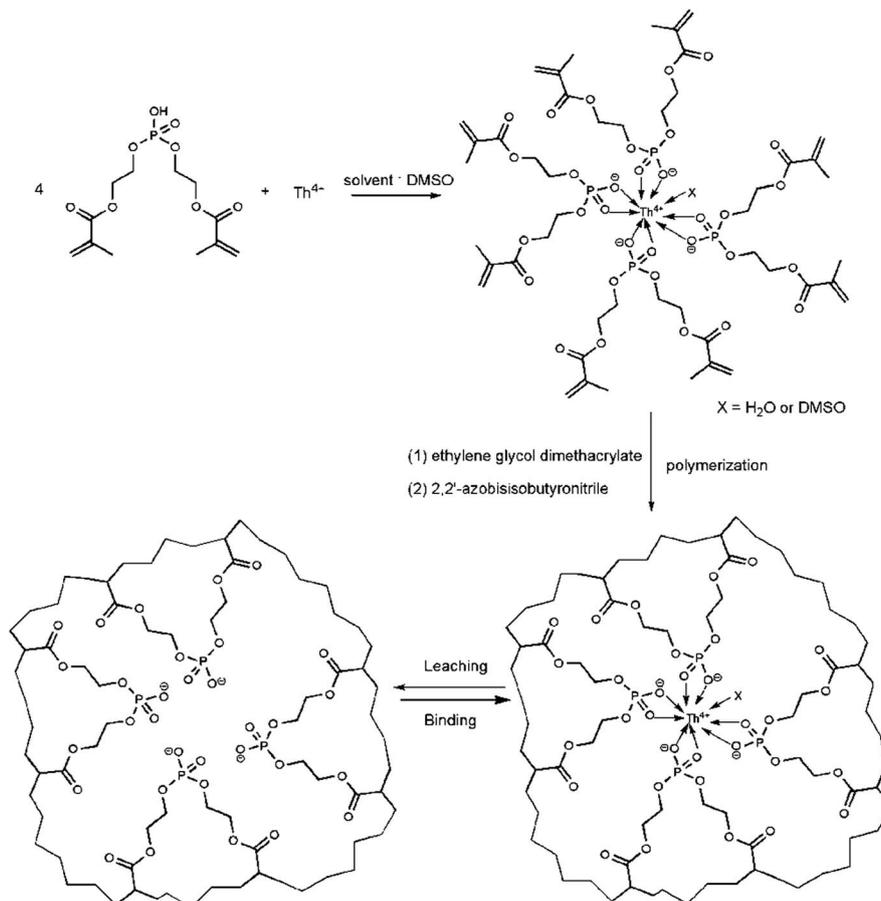
	CN	$R/\text{\AA}$	$\sigma^2/10^{-3} \text{\AA}^2$	$\Delta E_0/\text{eV}$	R factor
Th-O	9.1 ± 0.7	2.39 ± 0.01	5.2 ± 0.3	4.4 ± 1.0	0.004



Scheme 1 Scheme of the complex compound of DMP and Th^{4+} .

Then, fixing the molar ratio of EGDMA to BMAOP at 4, IIP1-IIP4, with the molar ratios of BMAOP to Th^{4+} at 3, 3.5, 4 and 5 respectively, were synthesized. Meanwhile, the synthesis process of IIP3 is exhibited in Scheme 2. As can be seen from Fig. 3, the adsorption capacity of IIP3 is the highest. However, NIP shows high adsorption capacity, too, which is slightly higher than that of IIP1. In general, imprinting effect or large surface area always leads to a high adsorption capacity. Herein, the surface areas of IIP1-4 and NIP were determined to be 164, 182, 217, 210, and 196 $\text{m}^2 \text{g}^{-1}$, respectively. For the larger surface area, the adsorption capacity of NIP is slightly higher than that of IIP1. As the imprinting effect, the adsorption capacity of IIP2 is higher than that of NIP although the surface area of IIP2 is slightly lower than that of NIP. The reasons for the large surface area and high adsorption capacity of NIP will be discussed in detail in Sections 3.2.3 and 3.3.3, respectively. We also digested the IIPs, and found that the residual thorium was less than 3 mg g^{-1} . In other words, the residual thorium in the IIPs has little influence on the adsorption capacities.

Furthermore, the selectivity coefficient ($\text{Th}^{4+}/\text{Eu}^{3+}$) of IIP3 are obviously higher than those of IIP1, IIP2, IIP4, and much higher than that of NIP (Fig. 3). This could be ascribed to the fact that one Th^{4+} ion coordinates with four BMAOP molecules, and excessive or insufficient functional ligands may reduce the



Scheme 2 Scheme for the preparation of Th^{4+} -IIP3.



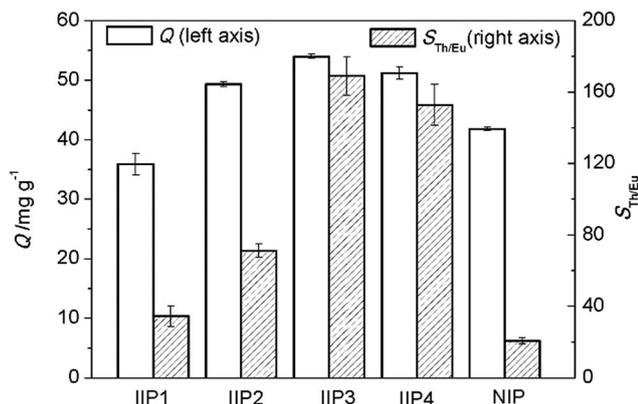


Fig. 3 Adsorption capacities and selectivity coefficients ($\text{Th}^{4+}/\text{Eu}^{3+}$) of IIPs with different ratios of BMAOP to Th^{4+} (from IIP1 to IIP4, $n(\text{BMAOP})/n(\text{Th}^{4+}) = 3, 3.5, 4,$ and 5) and NIP. (Adsorbent: 0.0050 g , metal ions concentration: 30 mg L^{-1} , solution: 0.0100 L , pH: 3.0 , time: 12 h).

specificity of the binding sites. The proper molar ratio of BMAOP to Th^{4+} should be $4 : 1$, consistent with the results of ^{31}P NMR, elemental analysis and EXAFS. Thus, IIP3 was selected in the following characterization and adsorption experiments.

3.2 Characterization

3.2.1 FT-IR spectra. The FT-IR spectra of the IIP3 and NIP (before and after washing with Na_2EDTA solution) show similar backbones because of their identical components (Fig. 4). All of the spectra show a strong absorption peak at 1725 cm^{-1} attributed to the stretching vibration of $\text{C}=\text{O}$ group and a peak at 1160 cm^{-1} assigned to the stretching vibration of $\text{P}=\text{O}$ group. These peaks demonstrate that BMAOP has been successfully copolymerized with EGDMA, and IIP3 has an identical composition as NIP. After washing, the absorption at 1022 cm^{-1} , the stretching vibration of $\text{S}=\text{O}$ group, diminished, indicating the leaching of the residual DMSO.

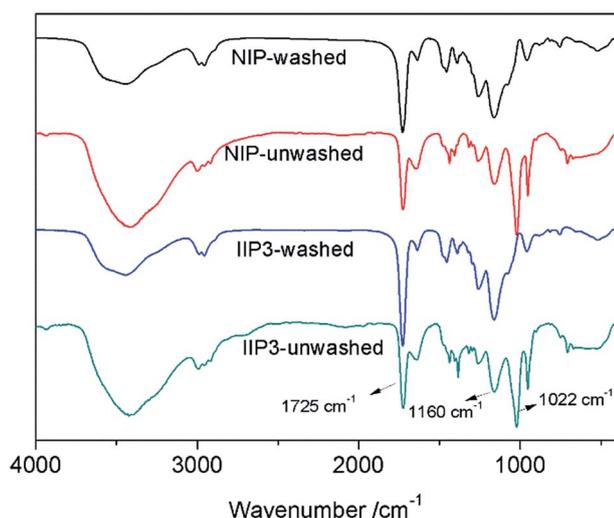


Fig. 4 FT-IR spectra of (un)washed IIP3 and NIP.

3.2.2 Thermal analysis (TGA). Thermal stabilities of the IIP3 and NIP particles were investigated by thermo-gravimetric analyses (Fig. 5). In the measurement process, the samples were heated from room temperature to $600 \text{ }^\circ\text{C}$ with a heating rate of $10 \text{ }^\circ\text{C min}^{-1}$ in air atmosphere. The TGA plots of IIP3 and NIP are very similar, indicating their identical components. Since the thermal decomposition temperatures of IIP3 and NIP are up to $240 \text{ }^\circ\text{C}$, they can be used stably in the common environments.

3.2.3 Surface area analysis. The surface area is an important parameter for adsorption materials which affect the adsorption kinetics and capacity greatly. Generally, it is believed that materials with larger surface area have a higher adsorption capacity.⁴⁰ Herein, the BET surface areas of the IIP3 and NIP were 217 and $196 \text{ m}^2 \text{ g}^{-1}$, respectively, and they had similar pore size (IIP3: 0.46 nm , NIP: 0.49 nm) and pore volume (IIP3: $8.4 \text{ cm}^3 \text{ g}^{-1}$, NIP: $9.9 \text{ cm}^3 \text{ g}^{-1}$). The large surface areas of IIP3 and NIP may be ascribed to the presence of DMSO in the preparation process. The functional ligand (BMAOP) has a similar methacryloxyethyl part as the cross-linker (EGDMA), leading to a similar polymerization ability. Since DMSO has a good thermodynamic compatibility for poly(EGDMA), the phase separation will be late in the polymerization, resulting in IIP3 (NIP) with large surface area.⁴¹

3.3 Adsorption properties

3.3.1 Effect of the acidity. The effect of acidity was investigated by altering the concentration of HCl from $10^{-3.8}$ ($\text{pH} = 3.8$) to 6.0 mol L^{-1} (Fig. 6). With the increase of the HCl concentration, the adsorption capacity of IIP3 decreases, while that of NIP increases to a maximum value at $\text{pH} = 2$ and then decreases. Moreover, the IIP3 exhibits a greater retention capacity than NIP at $\text{pH} < 1$ and $\text{pH} > 2$, while their retention capacities are almost identical at the pH range of $1\text{--}2$. This interesting phenomenon may be attributed to the competition adsorption. In acidic solution, H^+ competes with Th^{4+} for combining the phosphoric acid group on the polymer, which becomes much serious at $\text{pH} < 1$. When the pH value is higher than 2 , the Th^{4+} starts to hydrolyze, and OH^- competes with the

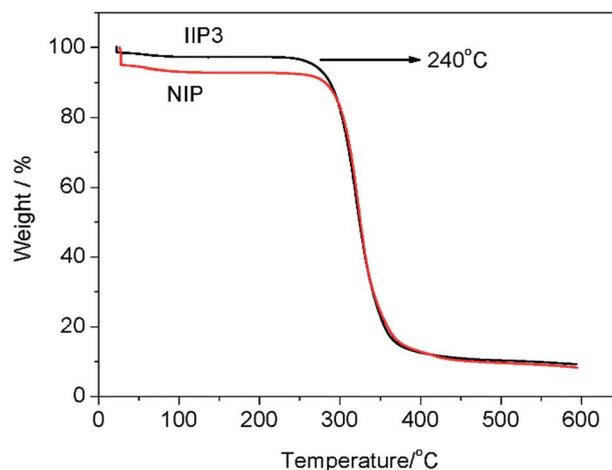


Fig. 5 TGA plots of IIP3 and NIP particles.



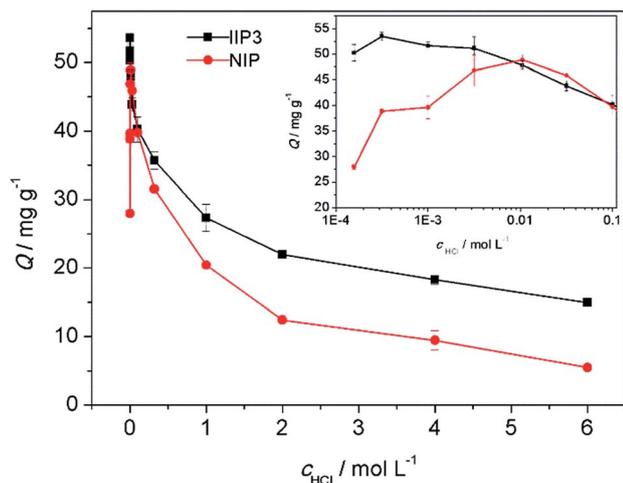


Fig. 6 Effect of acidity on the adsorption capacities of IIP3 and NIP. (Adsorbent: 0.0050 g, thorium concentration: 30 mg L⁻¹, thorium solution: 0.0100 L, time: 12 h).

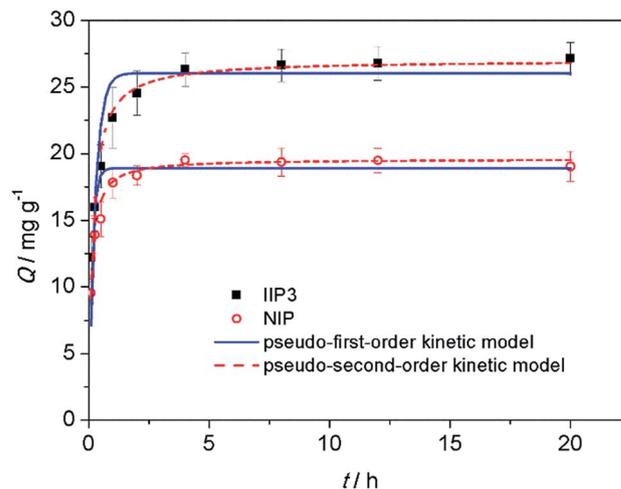


Fig. 7 Effect of time on the adsorption of Th⁴⁺ on the IIP3 and NIP. (Adsorbent: 0.0050 g, thorium concentration: 20 mg L⁻¹, thorium solution: 0.0100 L, HCl: 1.0 mol L⁻¹).

phosphoric acid group in the binding of Th⁴⁺. These two competitions lead to the fact that the adsorption capacity of NIP reaches its maximum value at pH = 2. While, in the intense competition situation, the specific binding sites in the IIP3 can combine with Th⁴⁺ more effectively than the randomly distributed functional groups in the NIP, leading to a larger adsorption capacity at pH < 1 and pH > 2. The IIP3 has a maximum adsorption capacity of 53.5 mg g⁻¹ at pH = 3.5, which is comparable with the theoretical capacity (52.1 mg g⁻¹) estimated according to the feed ratio. Different from the Th⁴⁺-IIPs reported in the literature which hardly adsorb Th⁴⁺ at pH = 1 (Table 1), the Th⁴⁺-IIP3 has 50% of the maximum adsorption capacity in 1.0 mol L⁻¹ HCl and still has 25% of the maximum adsorption capacity even in 6.0 mol L⁻¹ HCl. It may be attributed to the low pK_a value of the functional ligand, and the strong coordination interaction between Th⁴⁺ and phosphoric acid group. To investigate the adsorption ability of the Th⁴⁺-IIP3 in high acidity environment, the following experiments were performed in 1.0 mol L⁻¹ HCl.

3.3.2 Adsorption kinetics. According to the effect of time on the adsorption of Th⁴⁺ from aqueous solutions (Fig. 7), the equilibrium times of IIP3 and NIP were determined to be ca. 8 h. To guarantee adsorption equilibrium, the adsorption time was chosen to be 12 h in the following studies.

To explore the mechanism of adsorption kinetics, two different kinetic models were applied to fit the experimental data, *i.e.*, the pseudo-first-order kinetic model (eqn (5)) and the pseudo-second-order kinetic model (eqn (6)).

$$Q_t = Q_e(1 - e^{-k_1 t}) \quad (5)$$

$$Q_t = t / \left(\frac{t}{Q_e} + \frac{Q_e^2}{k_2} \right) \quad (6)$$

where Q_e (mg g⁻¹) and Q_t (mg g⁻¹) are the adsorption capacities at equilibrium and at time t (h), respectively; k_1 (h⁻¹) and k_2 (mg g⁻¹ h⁻¹) are the adsorption rate constants related to pseudo-

first-order and pseudo-second-order kinetic models, respectively. The pseudo-first-order kinetic model assumes that the rate of adsorption site occupation is proportional to the number of unoccupied sites, while the pseudo-second-order kinetic model corresponds to the chemical reaction mechanisms, in which the adsorption rate are controlled by chemical adsorption *via* sharing or exchange of electrons between the adsorbate and adsorbent.⁴² In this study, the pseudo-second-order kinetic model is much better than the pseudo-first-order kinetic model in the fitting of the experimental data (Fig. 7 and Table 3), suggesting that the adsorption is a chemical coordination process.

3.3.3 Adsorption isotherm. The adsorption capacities of IIP3 and NIP increased with the increase of equilibrium concentrations (Fig. 8). The isothermal adsorption data were analyzed with the Langmuir (eqn (7)) and Freundlich (eqn (8)) equations.

$$Q = Q_m b C_e / (1 + C_e) \quad (7)$$

$$Q = K_F C_e^{1/n} \quad (8)$$

where C_e (mg L⁻¹) is the equilibrium concentration of Th⁴⁺, Q_m (mg g⁻¹) is the maximum adsorption quantity, b (L mg⁻¹) is the Langmuir constant, K_F and n are the Freundlich constants. It can be seen from Table 4 and Fig. 8 that Langmuir model is more suitable in the fitting of the isothermal adsorption data of IIP3, while Freundlich model fits those of NIP better. The maximum adsorption capacity of IIP3 calculated from the Langmuir isotherm was 33.3 mg g⁻¹.

There are specific binding sites and unspecific binding sites in IIPs, while the binding sites in NIP is unspecific. The surface area of NIP is slightly lower than that of IIP3, suggesting that the accessibility of the binding sites of IIP3 and NIP is similar.

When the concentration of Th⁴⁺ is low, the specific binding sites in IIP3 show stronger binding ability than the unspecific binding sites in NIP. For example, at an initial Th⁴⁺



Table 3 Comparison of the fitting results of the pseudo-first-order and pseudo-second-order kinetic models

	$Q_e(\text{exp})/\text{mg g}^{-1}$	Pseudo-first-order			Pseudo-second-order		
		k_1/h^{-1}	R^2	$Q_e(\text{cal})/\text{mg g}^{-1}$	$k_2/\text{mg g}^{-1} \text{h}^{-1}$	R^2	$Q_e(\text{cal})/\text{mg g}^{-1}$
IIP3	26.8 ± 1	3.8 ± 0.5	0.87	26.0 ± 0.9	$(1.2 \pm 0.1) \times 10^5$	0.96	27.0 ± 0.5
NIP	19.3 ± 0.9	7.0 ± 1.0	0.89	18.9 ± 0.4	$(8.0 \pm 0.6) \times 10^4$	0.98	19.6 ± 0.2

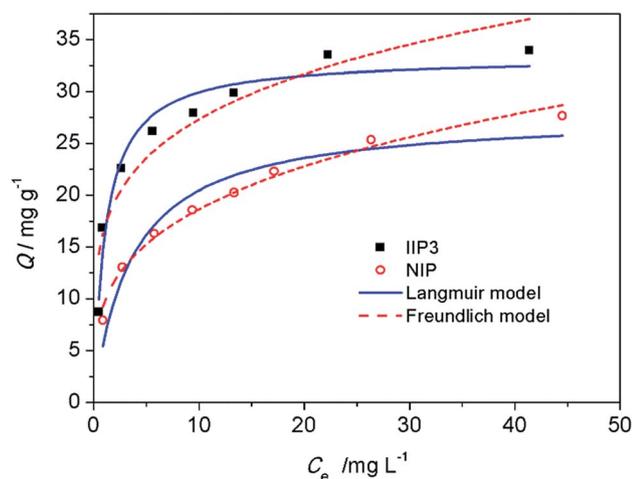
Fig. 8 Adsorption isotherms of IIP3 and NIP. (Adsorbent: 0.0050 g, thorium solution: 0.0100 L, HCl: 1.0 mol L⁻¹, time: 12 h).

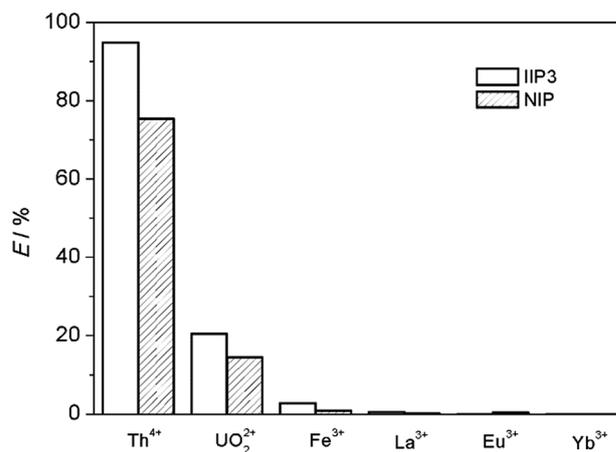
Table 4 Comparison of the fitting results of the Langmuir and Freundlich models

	Langmuir			Freundlich		
	R^2	$Q_m/\text{mg g}^{-1}$	$b/\text{L mg}^{-1}$	R^2	K_F	n
IIP3	0.95	33.3 ± 1.0	0.9 ± 0.2	0.88	17 ± 2	4.7 ± 0.7
NIP	0.93	27.8 ± 2.0	0.28 ± 0.06	0.99	9.6 ± 0.4	3.5 ± 0.2

concentration of 5.2 mg L⁻¹, the k_d values of IIP3 and NIP are 1.7×10^4 and 8.1×10^3 mL g⁻¹, respectively, indicating that the IIP3 can adsorb Th⁴⁺ more effectively. When the Th⁴⁺ in the solution is abundant, both the specific binding sites and unspecific binding sites in IIP3, as well as the unspecific binding sites in NIP can combine with Th⁴⁺. Thus, with the

increase of the equilibrium concentration of Th⁴⁺, the difference of adsorption capacities between IIP3 and NIP becomes smaller (Fig. 8).

3.3.4 Adsorption selectivity. To investigate the selectivity of the IIP3, competitive adsorptions were performed. In this investigation, UO₂²⁺, Fe³⁺, La³⁺, Eu³⁺, and Dy³⁺ were chosen as competitive metal ions because they often coexist in minerals as well as in waste water.^{3,4} The obtained k_d , S and k' (Table 5) show that IIP3 can effectively extract Th⁴⁺ from aqueous solution containing various interfering ions, especially in the presence of lanthanide ions. Also, 0.0100 g IIP3 (NIP) was used to extract Th⁴⁺ from a mixed solution of Th⁴⁺ and the above competing ions with each concentration of 20 mg L⁻¹. The result (Fig. 9) indicates that Th⁴⁺ can be effectively extracted from the mixture solution, and only a few UO₂²⁺ is co-extracted.

Fig. 9 Extraction of Th⁴⁺ and other competing ions from mixed solution. (Adsorbent: 0.0100 g, each metal ion concentration: 20 mg L⁻¹, solution: 0.0100 L, HCl: 1.0 mol L⁻¹, time: 12 h).Table 5 Selective adsorption properties of the IIP3 and NIP^a

Competing ions	$k_{d,\text{Th}} \times 10^{-3} (\text{mL g}^{-1})$		$k_{d,\text{M}} \times 10^{-3} (\text{mL g}^{-1})$		$S_{\text{Th/M}}$		k'
	IIP3	NIP	IIP3	NIP	IIP3	NIP	
UO ₂ ²⁺	38.7 ± 0.6	6.3 ± 0.5	0.33 ± 0.006	0.15 ± 0.003	116 ± 0.3	41 ± 2	2.9 ± 0.2
Fe ³⁺	27.8 ± 0.3	5.9 ± 0.1	—	—	—	—	—
La ³⁺	53 ± 3	7.3 ± 0.3	—	—	—	—	—
Eu ³⁺	62 ± 1	7.1 ± 0.1	—	—	—	—	—
Yb ³⁺	52 ± 3	7.2 ± 0.1	—	—	—	—	—

^a The adsorptions of Fe³⁺ and Ln³⁺ on the IIP3 and NIP could not be detected.

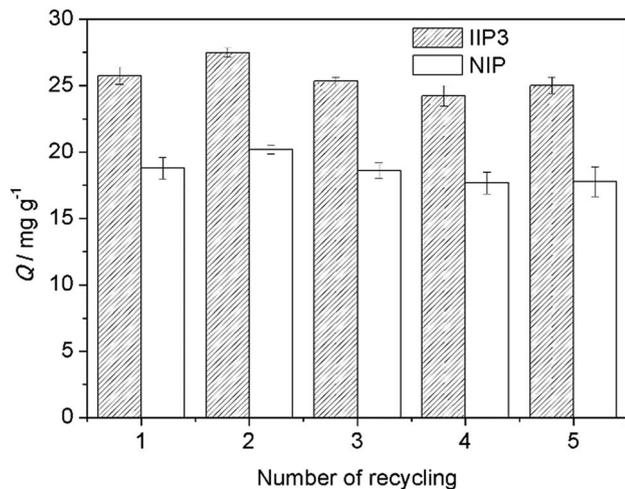


Fig. 10 The retention capacities of the recycled Th⁴⁺-IIP3 and NIP particles in five times. (Adsorbent: 0.0050 g, thorium concentration: 20 mg L⁻¹, thorium solution: 0.0100 L, HCl: 1.0 mol L⁻¹, time: 12 h).

3.3.5 Reusability. The reusability is one of the important performances of IIPs, affecting their practical application. To test the reusability of the polymers, subsequent sorption and desorption cycles were performed using 0.0050 g IIP3 (NIP), in which 10 mL 20 mg L⁻¹ Th⁴⁺ in 1.0 mol L⁻¹ HCl was used in the adsorption process and 10 mL 0.1 mol L⁻¹ Na₂EDTA was applied in the leaching experiment. As shown in Fig. 10, the adsorption capacities of the IIP3 and NIP hardly decrease after five cycles, indicating that the Th⁴⁺-IIP3 (NIP) has good reusability in high acidity environment.

4. Conclusions

For the first time, a new Th⁴⁺-IIP that could be used in high acidity environment was synthesized using BMAOP as functional ligand. The molar ratio of the functional ligand to Th⁴⁺ was optimized to be 4 by ³¹P NMR titration, elemental analysis, and EXAFS, which was verified by comparing the adsorption capacities and selectivities of Th⁴⁺-IIPs with different compositions. The best Th⁴⁺-IIP3, with the molar ratio of BMAOP to Th⁴⁺ at 4, had a maximum adsorption capacity of 33.3 mg g⁻¹ in 1 mol L⁻¹ HCl solution, and its adsorption capacity was still considerable even in 6.0 mol L⁻¹ HCl. The adsorption kinetic of the IIP3 followed the pseudo-second-order kinetic model. And the IIP3 can effectively extract Th⁴⁺ from the mixed aqueous solution of Th⁴⁺, La³⁺, Eu³⁺, Yb³⁺, UO₂²⁺, and Fe³⁺. Moreover, the IIP3 can be reused for at least five times without obvious loss of adsorption capacity. Therefore, the prepared Th⁴⁺-IIP3 could be used in the detection and removal of Th⁴⁺ in high acidity waste water. Future studies will be focused on the improvement of the adsorption capacity and the selectivity coefficient (Th⁴⁺/UO₂²⁺) in high acidity environment for the treatment of the real samples.

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Notes and references

- S. C. Zhang, P. Liu and B. J. Zhang, *World Nucl. Geosci.*, 2005, **22**, 98–103.
- Z. Gu, *Chin. J. Nucl. Sci. Eng.*, 2007, **27**, 97–105.
- J. Cheng, Y. Hou and L. Che, *Chin. Rare Earths*, 2008, **29**, 76–77.
- Z. Zhu, Y. Pranolo and C. Y. Cheng, *Miner. Eng.*, 2015, **77**, 185–196.
- I. Shtangeeva, S. Ayrault and J. Jain, *J. Environ. Radioact.*, 2005, **81**, 283–293.
- J. Fu, Q. D. Chen, T. X. Sun and X. H. Shen, *Sep. Purif. Technol.*, 2013, **119**, 66–71.
- S. Kesava Raju Ch and M. S. Subramanian, *J. Hazard. Mater.*, 2007, **145**, 315–322.
- T. P. Rao, P. Metilda and J. M. Gladis, *Talanta*, 2006, **68**, 1047–1064.
- R. Pilviö and M. Bickel, *J. Alloys Compd.*, 1998, **271–273**, 49–53.
- D. Nanda, M. S. Oak, M. P. Kumar, B. Maiti and P. K. Dutta, *Sep. Sci. Technol.*, 2001, **36**, 2489–2497.
- W. A. Wan Ibrahim, L. I. Abd Ali, A. Sulaiman, M. M. Sanagi and H. Y. Aboul-Enein, *Crit. Rev. Anal. Chem.*, 2014, **44**, 233–254.
- C. Branger, W. Meouche and A. Margailan, *React. Funct. Polym.*, 2013, **73**, 859–875.
- J. Fu, L. Chen, J. Li and Z. Zhang, *J. Mater. Chem. A*, 2015, **3**, 13598–13627.
- H. L. Liang, Q. D. Chen and X. H. Shen, *J. Nucl. Radiochem. Sci.*, 2016, **38**, 129–144.
- H. He, Q. Gan and C. Feng, *RSC Adv.*, 2017, **7**, 15102–15111.
- Z. Zhang, J. Li, X. Song, J. Ma and L. Chen, *RSC Adv.*, 2014, **4**, 46444–46453.
- F. F. He, H. Q. Wang, Y. Y. Wang, X. F. Wang, H. S. Zhang, H. L. Li and J. H. Tang, *J. Radioanal. Nucl. Chem.*, 2012, **295**, 167–177.
- A.-S. Chauvin, J.-C. G. Bünzli, F. Bochud, R. Scopelliti and P. Froidevaux, *Chem.–Eur. J.*, 2006, **12**, 6852–6864.
- M. Monier and N. H. Elsayed, *J. Colloid Interface Sci.*, 2014, **423**, 113–122.
- J. Fasihi, S. Ammari Alahyari, M. Shamsipur, H. Sharghi and A. Charkhi, *React. Funct. Polym.*, 2011, **71**, 803–808.
- S. Buyuktiryaki, R. Say, A. Ersoz, E. Birlik and A. Denizli, *Talanta*, 2005, **67**, 640–645.
- C. Lin, H. Wang, Y. Wang and Z. Cheng, *Talanta*, 2010, **81**, 30–36.
- G. Huang, Z. Chen, L. Wang, T. Lv and J. Shi, *J. Radioanal. Nucl. Chem.*, 2016, **310**, 1265–1272.
- J. Fasihi, N. Bavarsad, S. Shariati and K. Ashtari, *Int. J. Environ. Anal. Chem.*, 2016, **96**, 789–800.
- E. Birlik, S. Buyuktiryaki, A. Ersoz, A. Denizli and R. Say, *Sep. Sci. Technol.*, 2006, **41**, 3109–3121.
- Q. He, X. Chang, Q. Wu, X. Huang, Z. Hu and Y. Zhai, *Anal. Chim. Acta*, 2007, **605**, 192–197.



- 27 C. R. Lin, H. Q. Wang, Y. Y. Wang, L. Zhou and J. Liang, *Int. J. Environ. Anal. Chem.*, 2011, **91**, 1050–1061.
- 28 Z. Cheng, H. Wang, Y. Wang, F. He, H. Zhang and S. Yang, *Microchim. Acta*, 2011, **173**, 423–431.
- 29 S. Mishra and D. K. Singh, *Desalin. Water Treat.*, 2014, **56**, 1364–1371.
- 30 G. X. Xu, *Rare earths*, Metallurgical Industry Press, Beijing, China, 2nd edn, 1995.
- 31 H. Tong, Y. Wang, W. Liao and D. Li, *Sep. Purif. Technol.*, 2013, **118**, 487–491.
- 32 S. Gao, T. X. Sun, Q. D. Chen and X. H. Shen, *Radiochim. Acta*, 2016, **104**, 457–469.
- 33 S. Gao, X. H. Shen, Q. D. Chen and H. C. Gao, *Sci. China: Chem.*, 2012, **55**, 1712–1718.
- 34 K. Uezu, T. Kuwabara, M. Yoshida and G. Masahiro, *Anal. Sci.*, 2004, **20**, 1593–1597.
- 35 M. Yoshida, K. Uezu, M. Goto and S. Furusaki, *Macromolecules*, 1999, **32**, 1237–1243.
- 36 F. H. El-Sweify, A. A. Abdel-Fattah and S. M. Ali, *J. Chem. Thermodyn.*, 2008, **40**, 798–805.
- 37 E. V. Goud, D. Das, A. Sivaramakrishna, K. Vijayakrishna, V. Sabareesh, G. Gopakumar, C. V. S. Brahmmananda Rao, M. Y. Lone and P. C. Jha, *Polyhedron*, 2016, **117**, 741–748.
- 38 H. Yu, X. Wei, J. Li, S. Gu, S. Zhang, L. Wang, J. Ma, L. Li, Q. Gao, R. Si, F. Sun, Y. Wang, F. Song, H. Xu, X. Yu, Y. Zou, J. Wang, Z. Jiang and Y. Huang, *Nucl. Sci. Tech.*, 2015, **26**, 050102.
- 39 M. Newville, *J. Synchrotron Radiat.*, 2001, **8**, 322–324.
- 40 P. Metilda, J. Mary Gladis and T. Prasada Rao, *Anal. Chim. Acta*, 2004, **512**, 63–73.
- 41 W. Meouche, K. Laatikainen, A. Margaillan, T. Silvonon, H. Siren, T. Sainio, I. Beurroies, R. Denoyel and C. Branger, *Eur. Polym. J.*, 2017, **87**, 124–135.
- 42 A. R. Iftikhar, H. N. Bhatti, M. A. Hanif and R. Nadeem, *J. Hazard. Mater.*, 2009, **161**, 941–947.

